

**TEST PAPER OF JEE(MAIN) EXAMINATION – 2019**  
**(Held On Friday 11<sup>th</sup> JANUARY, 2019) TIME : 02 : 30 PM To 05 : 30 PM**  
**CHEMISTRY**

1. The correct option with respect to the Pauling electronegativity values of the elements is :-  
(1) Ga < Ge                      (2) Si < Al  
(3) P > S                          (4) Te > Se

Ans. (1)

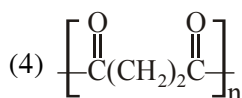
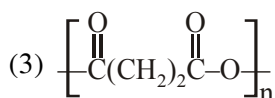
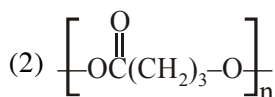
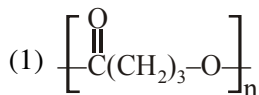
Sol.

B      C  
Al     Si

Ga < Ge

Along the period electronegativity increases

2. The homopolymer formed from 4-hydroxybutanoic acid is :-



Ans. (1)

Sol.



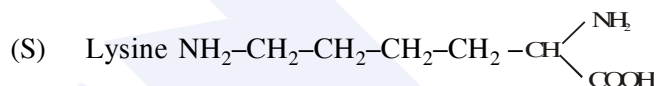
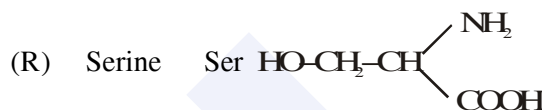
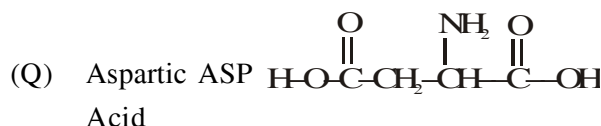
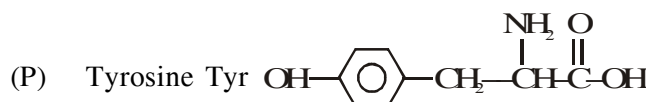
3. The correct match between Item I and Item II is :-

Item I		Item II	
(A)	Ester test	(P)	Tyr
(B)	Carbylamine test	(Q)	Asp
(C)	Phthalein dye test	(R)	Ser
		(S)	Lys

- (1) (A)→(Q); (B)→(S); (C)→(P)  
(2) (A)→(R); (B)→(Q); (C)→(P)  
(3) (A)→(Q); (B)→(S); (C)→(R)  
(4) (A)→(R); (B)→(S); (C)→(Q)

Ans. (1)

Sol.



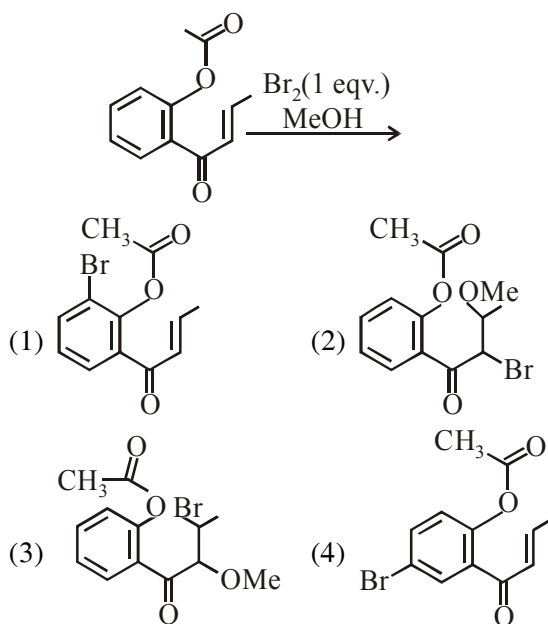
- (A) Ester test      (Q) Aspartic acid (Acidic amino acid)  
(B) Carbylamine test      (S) Lysine [NH<sub>2</sub> group present]  
(C) Phthalein dye      (P) Tyrosine {Phenolic group present}

4. Taj Mahal is being slowly disfigured and discoloured. This is primarily due to :-  
(1) Water pollution      (2) Global warming  
(3) Soil pollution      (4) Acid rain

Ans. (4)

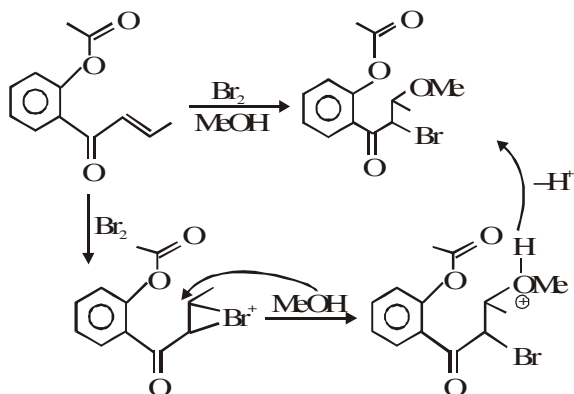
Sol. Taj mahal is slowly disfigured and discoloured due to acid rain.

5. The major product obtained in the following conversion is :-



Ans. (2)

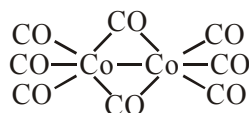
Sol.



6. The number of bridging CO ligand (s) and Co-Co bond (s) in  $\text{Co}_2(\text{CO})_8$ , respectively are :-  
 (1) 0 and 2                      (2) 2 and 0  
 (3) 4 and 0                      (4) 2 and 1

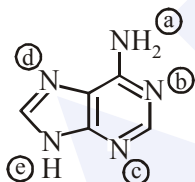
Ans. (4)

Sol.



Bridging CO are 2 and Co - Co bond is 1.

7. In the following compound,



the favourable site/s for protonation is/are :-

- (1) (b), (c) and (d)              (2) (a)  
 (3) (a) and (e)                      (4) (a) and (d)

Ans. (1)

Sol. Localised lone pair  $e^-$ .

8. The higher concentration of which gas in air can cause stiffness of flower buds ?  
 (1)  $\text{SO}_2$                               (2)  $\text{NO}_2$   
 (3)  $\text{CO}_2$                               (4) CO

Ans. (1)

Sol. Due to acid rain in plants high concentration of  $\text{SO}_2$  makes the flower buds stiff and makes them fall.

9. The correct match between item I and item II is :-

Item I		Item II	
(A)	Allosteric effect	(P)	Molecule binding to the active site of enzyme
(B)	Competitive inhibitor	(Q)	Molecule crucial for communication in the body
(C)	Receptor	(R)	Molecule binding to a site other than the active site of enzyme
(D)	Poison	(S)	Molecule binding to the enzyme covalently

- (1) (A)→(P); (B)→(R); (C)→(S); (D)→(Q)  
 (2) (A)→(R); (B)→(P); (C)→(S); (D)→(Q)  
 (3) (A)→(P); (B)→(R); (C)→(Q); (D)→(S)  
 (4) (A)→(R); (B)→(P); (C)→(Q); (D)→(S)

Ans. (4)

10. The radius of the largest sphere which fits properly at the centre of the edge of body centred cubic unit cell is : (Edge length is represented by 'a') :-

- (1) 0.134 a                              (2) 0.027 a  
 (3) 0.067 a                              (4) 0.047 a

Ans. (3)

Sol.

$$a = 2(R + r)$$

$$\frac{a}{2} = (R + r) \dots(1)$$

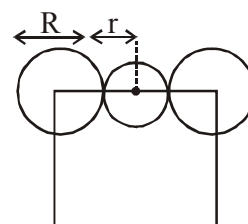
$$a\sqrt{3} = 4R \dots(2)$$

Using (1) & (2)

$$\frac{a}{2} = \frac{a\sqrt{3}}{4} = r$$

$$a \left( \frac{2 - \sqrt{3}}{4} \right) = r$$

$$r = 0.067 a$$



11. Among the colloids cheese (C), milk (M) and smoke (S), the correct combination of the dispersed phase and dispersion medium, respectively is :-

- (1) C : solid in liquid; M : solid in liquid; S : solid in gas
- (2) C : solid in liquid; M : liquid in liquid; S : gas in solid
- (3) C : liquid in solid; M : liquid in solid; S : solid in gas
- (4) C : liquid in solid; M : liquid in liquid; S : solid in gas

Ans. (4)

Sol.

	Dispersed Phase	Dispersion Medium
Cheese	Liquid	Solid
Milk	Liquid	Liquid
Smoke	Solid	Gas

12. The reaction that does NOT define calcination is:-

- (1)  $\text{ZnCO}_3 \xrightarrow{\Delta} \text{ZnO} + \text{CO}_2$
- (2)  $\text{Fe}_2\text{O}_3 \cdot \text{XH}_2\text{O} \xrightarrow{\Delta} \text{Fe}_2\text{O}_3 + \text{XH}_2\text{O}$
- (3)  $\text{CaCO}_3 \cdot \text{MgCO}_3 \xrightarrow{\Delta} \text{CaO} + \text{MgO} + 2 \text{CO}_2$
- (4)  $2 \text{Cu}_2\text{S} + 3 \text{O}_2 \xrightarrow{\Delta} 2 \text{Cu}_2\text{O} + 2 \text{SO}_2$

Ans. (4)

Sol. Calcination is carried out for carbonates and oxide ores in absence of oxygen. Roasting is carried out mainly for sulphide ores in presence of excess of oxygen.

13. The reaction,

$\text{MgO(s)} + \text{C(s)} \rightarrow \text{Mg(s)} + \text{CO(g)}$ , for which  $\Delta_r H^\circ = + 491.1 \text{ kJ mol}^{-1}$  and  $\Delta_r S^\circ = 198.0 \text{ JK}^{-1} \text{ mol}^{-1}$ , is not feasible at 298 K. Temperature above which reaction will be feasible is :-

- (1) 1890.0 K
- (2) 2480.3 K
- (3) 2040.5 K
- (4) 2380.5 K

Ans. (2)

Sol.  $T_{\text{eq}} = \frac{\Delta H}{\Delta S}$

$$= \frac{491.1 \times 1000}{198}$$

$$= 2480.3 \text{ K}$$

14. Given the equilibrium constant :

KC of the reaction :

$\text{Cu(s)} + 2\text{Ag}^+(\text{aq}) \rightarrow \text{Cu}^{2+}(\text{aq}) + 2\text{Ag(s)}$  is  $10 \times 10^{15}$ , calculate the  $E_{\text{cell}}^0$  of this reaction at 298 K

$$\left[ 2.303 \frac{RT}{F} \text{ at } 298 \text{ K} = 0.059 \text{ V} \right]$$

- (1) 0.04736 V
- (2) 0.4736 V
- (3) 0.4736 mV
- (4) 0.04736 mV

Ans. (2)

Sol.  $E_{\text{cell}} = E_{\text{cell}}^0 - \frac{0.059}{n} \log Q$

At equilibrium

$$E_{\text{Cell}}^0 = \frac{0.059}{2} \log 10^{16}$$

$$= 0.059 \times 8$$

$$= 0.472 \text{ V}$$

15. The hydride that is NOT electron deficient is:-

- (1)  $\text{B}_2\text{H}_6$
- (2)  $\text{AlH}_3$
- (3)  $\text{SiH}_4$
- (4)  $\text{GaH}_3$

Ans. (3)

Sol. (1)  $\text{B}_2\text{H}_6$  : Electron deficient

(2)  $\text{AlH}_3$  : Electron deficient

(3)  $\text{SiH}_4$  : Electron precise

(4)  $\text{GaH}_3$  : Electron deficient

**MAJOR COMPUTER BASED TEST (CBT) SERIES**

**JEE (Advanced)- Target 2019**

[dlp.allen.ac.in](http://dlp.allen.ac.in)

Test Dates: 3<sup>rd</sup> Feb, 21<sup>st</sup> & 28<sup>th</sup> April, 12<sup>th</sup> May

0744-2750275

16. The standard reaction Gibbs energy for a chemical reaction at an absolute temperature T is given by

$$\Delta_r G^\circ = A - Bt$$

Where A and B are non-zero constants. Which of the following is TRUE about this reaction ?

- (1) Exothermic if  $B < 0$
- (2) Exothermic if  $A > 0$  and  $B < 0$
- (3) Endothermic if  $A < 0$  and  $B > 0$
- (4) Endothermic if  $A > 0$

Ans. (4)

Sol. Theory

17.  $K_2HgI_4$  is 40% ionised in aqueous solution. The value of its van't Hoff factor (i) is :-

- (1) 1.8
- (2) 2.2
- (3) 2.0
- (4) 1.6

Ans. (1)

Sol. For  $K_2[HgI_4]$   
 $i = 1 + 0.4(3-1)$   
 $= 1.8$

18. The de Broglie wavelength ( $\lambda$ ) associated with a photoelectron varies with the frequency ( $\nu$ ) of the incident radiation as, [ $\nu_0$  is threshold frequency] :

- (1)  $\lambda \propto \frac{1}{(\nu - \nu_0)^{\frac{3}{2}}}$
- (2)  $\lambda \propto \frac{1}{(\nu - \nu_0)^{\frac{1}{2}}}$
- (3)  $\lambda \propto \frac{1}{(\nu - \nu_0)^{\frac{1}{4}}}$
- (4)  $\lambda \propto \frac{1}{(\nu - \nu_0)}$

Ans. (2)

Sol. For electron

$$\lambda_{DB} = \frac{h}{\sqrt{2mK.E.}} \quad (\text{de broglie wavelength})$$

By photoelectric effect

$$h\nu = h\nu_0 + KE$$

$$KE = h\nu - h\nu_0$$

$$\lambda_{DB} = \frac{h}{\sqrt{2m \times (h\nu - h\nu_0)}}$$

$$\lambda_{DB} \propto \frac{1}{(\nu - \nu_0)^{\frac{1}{2}}}$$

19. The reaction  $2X \rightarrow B$  is a zeroth order reaction. If the initial concentration of X is 0.2 M, the half-life is 6 h. When the initial concentration of X is 0.5 M, the time required to reach its final concentration of 0.2 M will be :-

- (1) 18.0 h
- (2) 7.2 h
- (3) 9.0 h
- (4) 12.0 h

Ans. (1)

Sol. For zero order

$$[A_0] - [A_t] = kt$$

$$0.2 - 0.1 = k \times 6$$

$$k = \frac{1}{60} \text{ M/hr}$$

$$\text{and } 0.5 - 0.2 = \frac{1}{60} \times t$$

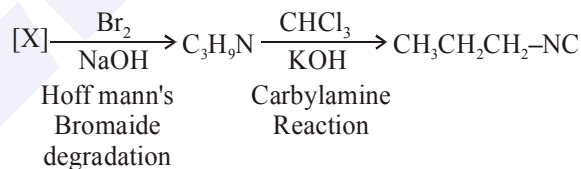
$$t = 18 \text{ hrs.}$$

20. A compound 'X' on treatment with  $Br_2/NaOH$ , provided  $C_3H_9N$ , which gives positive carbylamine test. Compound 'X' is :-

- (1)  $CH_3COCH_2NHCH_3$
- (2)  $CH_3CH_2COCH_2NH_2$
- (3)  $CH_3CH_2CH_2CONH_2$
- (4)  $CH_3CON(CH_3)_2$

Ans. (3)

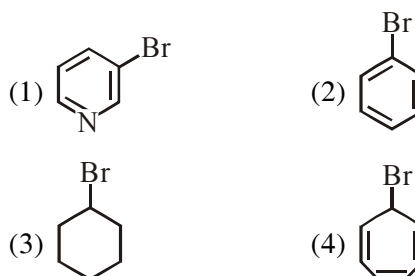
Sol.



Thus [X] must be aride with oen carbon more than is amine.

Thus [X] is  $CH_3CH_2CH_2CONH_2$

21. Which of the following compounds will produce a precipitate with  $AgNO_3$  ?



Ans. (4)

**MAJOR COMPUTER BASED TEST (CBT) SERIES**

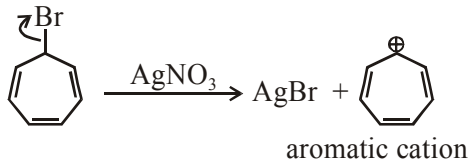
**JEE (Main)- Target 2019**

[dlp.allen.ac.in](http://dlp.allen.ac.in)

Test Dates: 24<sup>th</sup> & 31<sup>st</sup> March

0744-2750275

Sol.



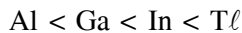
as it can produce aromatic cation so will produce precipitate with  $\text{AgNO}_3$ .

22. The relative stability of +1 oxidation state of group 13 elements follows the order :-

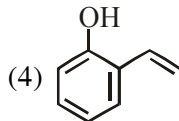
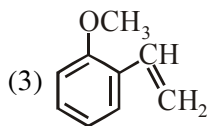
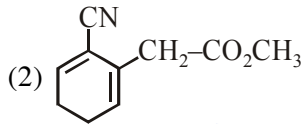
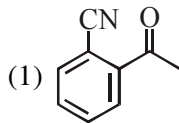
- (1)  $\text{Al} < \text{Ga} < \text{Tl} < \text{In}$  (2)  $\text{Tl} < \text{In} < \text{Ga} < \text{Al}$   
 (3)  $\text{Al} < \text{Ga} < \text{In} < \text{Tl}$  (4)  $\text{Ga} < \text{Al} < \text{In} < \text{Tl}$

Ans. (3)

Sol. Due to inert pair effect as we move down the group in 13<sup>th</sup> group lower oxidation state becomes more stable.

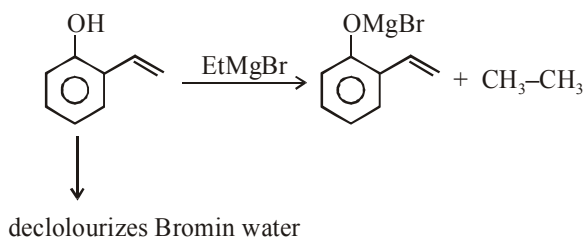


23. Which of the following compounds reacts with ethylmagnesium bromide and also decolourizes bromine water solution :-



Ans. (4)

Sol.



24. Match the following items in column I with the corresponding items in column II.

Column I		Column II	
(i)	$\text{Na}_2\text{CO}_3 \cdot 10 \text{H}_2\text{O}$	(P)	Portland cement ingredient
(ii)	$\text{Mg}(\text{HCO}_3)_2$	(Q)	Castner-Keller process
(iii)	$\text{NaOH}$	(R)	Solvay process
(iv)	$\text{Ca}_3\text{Al}_2\text{O}_6$	(S)	Temporary hardness

(1) (i)→(C); (ii)→(B); (iii)→(D); (iv)→(A)

(2) (i)→(C); (ii)→(D); (iii)→(B); (iv)→(A)

(3) (i)→(D); (ii)→(A); (iii)→(B); (iv)→(C)

(4) (i)→(B); (ii)→(C); (iii)→(A); (iv)→(D)

Ans. (2)

Sol.  $\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O} \rightarrow$  Solvay process

$\text{Mg}(\text{HCO}_3)_2 \rightarrow$  Temporary hardness

$\text{NaOH} \rightarrow$  Castner-kellner cell

$\text{Ca}_3\text{Al}_2\text{O}_6 \rightarrow$  Portland cement

25. 25 ml of the given HCl solution requires 30 mL of 0.1 M sodium carbonate solution. What is the volume of this HCl solution required to titrate 30 mL of 0.2 M aqueous NaOH solution?

(1) 25 mL (2) 50 mL (3) 12.5 mL (4) 75 mL

Ans. (1)

Sol. HCl with  $\text{Na}_2\text{CO}_3$

Eq. of HCl = Eq. of  $\text{Na}_2\text{CO}_3$

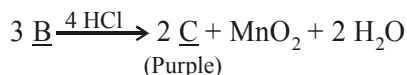
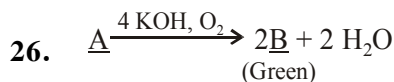
$$\frac{25}{1000} \times M \times 1 = \frac{30}{1000} \times 0.1 \times 2$$

$$M = \frac{6}{25} M$$

Eq of HCl = Eq. of NaOH

$$\frac{6}{25} \times 1 \times \frac{V}{1000} = \frac{30}{1000} \times 0.2 \times 1$$

$$V = 25 \text{ ml}$$

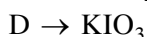
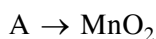
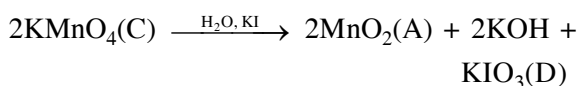
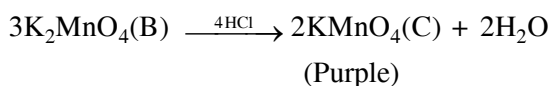
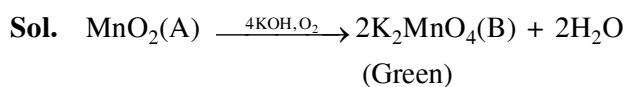


In the above sequence of reactions,

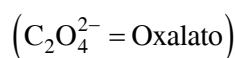
A and D respectively, are :-

- (1)  $\text{KIO}_3$  and  $\text{MnO}_2$       (2)  $\text{KI}$  and  $\text{K}_2\text{MnO}_4$   
(3)  $\text{MnO}_2$  and  $\text{KIO}_3$       (4)  $\text{KI}$  and  $\text{KMnO}_4$

Ans. (3)



27. The coordination number of Th in  $\text{K}_4[\text{Th}(\text{C}_2\text{O}_4)_4(\text{OH}_2)_2]$  is :-



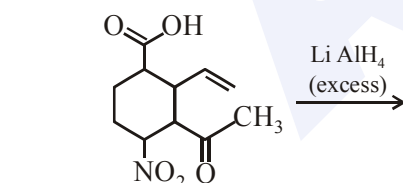
- (1) 6      (2) 10      (3) 14      (4) 8

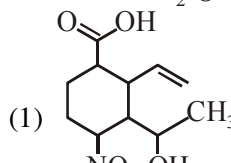
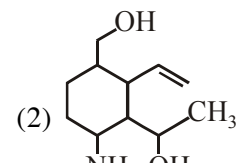
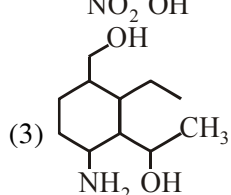
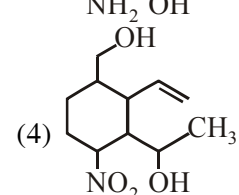
Ans. (2)

Sol.  $\text{C}_2\text{O}_4^{2-}$  (oxalato) : bidentate

$\text{H}_2\text{O}$  (aqua) : Monodentate

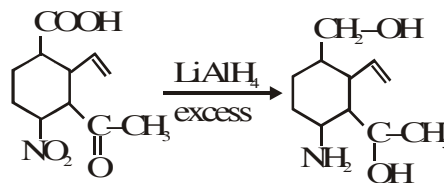
28. The major product obtained in the following reaction is :-



- (1)       (2)   
(3)       (4) 

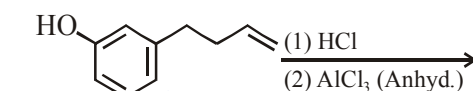
Ans. (2)

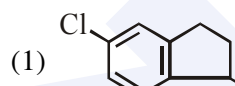
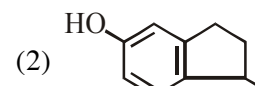
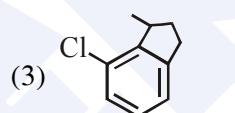
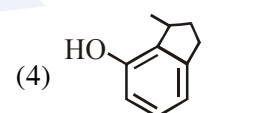
Sol.



$\text{LiAlH}_4$  will not affect  $\text{C}=\text{C}$  in this compound.

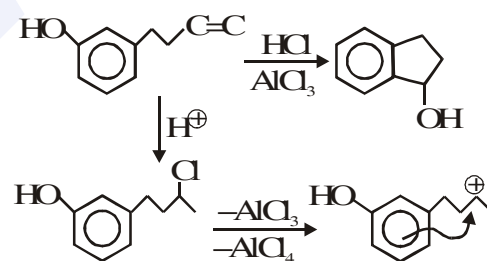
29. The major product of the following reaction is :-



- (1)       (2)   
(3)       (4) 

Ans. (2)

Sol.



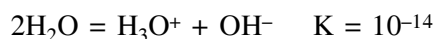
30. For the equilibrium,

$2\text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ , the value of  $\Delta G^\circ$  at 298 K is approximately :-

- (1)  $-80 \text{ kJ mol}^{-1}$       (2)  $-100 \text{ kJ mol}^{-1}$   
(3)  $100 \text{ kJ mol}^{-1}$       (4)  $80 \text{ kJ mol}^{-1}$

Ans. (4)

Sol.



$\Delta G^\circ = -RT \ln K$

$= \frac{-8.314}{1000} \times 298 \times \ln 10^{-14}$

$= 80 \text{ KJ/Mole}$